**Oxidation Numbers (charges):**



**Ionic Compounds -** one metal and one non-metal

**Steps to Writing Chemical Formulas:**

1. Write positive ion first
2. Write negative ion second
3. Add the oxidation numbers
	1. If they equal zero, then you are done
	2. If they don’t equal zero, then figure out how many of each you need (“making the charges balance” or “Criss-Cross Method”)
* move the charge of one atom and make it the subscript for the other

**Practice Compounds to Create:**

1. K and Se \_\_\_\_\_\_\_\_\_\_\_\_

2. Na and OH \_\_\_\_\_\_\_\_\_\_\_\_

3. NH4 and S \_\_\_\_\_\_\_\_\_\_\_\_

4. C2H3O2 and Li \_\_\_\_\_\_\_\_\_\_\_

5. CO3 and Ga \_\_\_\_\_\_\_\_\_\_\_\_

6. PO4 and Na \_\_\_\_\_\_\_\_\_\_\_\_